



I Can Statement - Year 8

Reactivity Series

Unit/Topic Overview	Pre-taught Vocabulary	
The reactivity series is a chart that ranks metals from the most reactive to the least reactive based on how easily they take part in chemical reactions. At the top of the list, highly reactive metals like potassium and sodium react vigorously with water and acids, quickly forming new compounds. In contrast, unreactive metals like gold and platinum sit at the very bottom, remaining completely unchanged over time. Chemists use this series to predict displacement reactions, where a more reactive metal kicks out a less reactive metal from its compound. It also helps us understand how metals are extracted from the Earth, as more reactive metals bond tightly to rocks and require extra energy to break free.	reactivity	Native state
	displacement	ore
	extraction	Blast furnace
	vigorous	electrolysis
	unreactive	compound

RAG Rating:		Start	End
1.	I am familiar with periodic table and can discuss trends in groups and periods.		
2.	I can describe alkali metals and their relative reactivity in detail		
3.	I can describe the reaction of metal + acid to give salt + hydrogen Students know test for Hydrogen gas is squeaky pop		
4.	I can describe the reaction between metal carbonate and acid, predicting products, and know test for carbon dioxide is limewater turns cloudy		
5.	I can describe the reaction of metal oxides and acid, predicting reactions from given chemicals. Students know test for oxygen is relighting a glowing splint		
6.	I can arrange metals in order of reactivity through observation		
7.	I can explain displacement reactions using the reactivity series		
8.	I understand that metals may be found as elements or as ore and may be extracted using carbon in a blast furnace		